

Net Ionic Equation Yahoo Answers

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Net Ionic Equation Yahoo Answers Please help with net ionic equations! If you can do only a few I'm good with that too. Thanks! Ammonium carbonate and Ba^{2+} Ammonium carbonate and Ca^{2+} Na_2HPO_4 and Mg^{2+} Oxidation of manganese Formation of BaCO_3 Formation of CaCO_3 Reaction of Na_2SO_4 and Ba^{2+} Reaction of $(\text{NH}_4)_2\text{C}_2\text{O}_4$ and Ca^{2+} potassium iodide and Cu^{2+} Net ionic equations? | Yahoo Answers What's the net ionic equation for: 1. aluminum iodide and chromium(III) acetate 2. ammonium nitrate and chromium(II) chloride Net ionic equations? | Yahoo Answers 1. What is the net ionic equation of the

reaction of FeCl₂ with NaOH? 2. What is the net ionic equation of the reaction of MgSO₄ with Pb(NO₃)₂? Express your answer as a chemical equation including phases. How do you find the net ionic equation? I'm having a hard time on these two particular questions. I was reading and it showed me that I need to do something with the charges but I could not ... Net ionic equations? | Yahoo Answers I'm stuck on this question. I need to find the net ionic equation for this reaction.
$$\text{Hg}_2\text{Cl}_2 (\text{s}) + 2\text{NH}_3 (\text{aq}) \rightarrow \text{Hg} (\text{s}) + \text{HgNH}_2\text{Cl} (\text{s}) + \text{NH}_4\text{Cl} (\text{aq})$$
 I got this for the total ionic equation :
$$\text{Hg}_2\text{Cl}_2 (\text{s}) + 2\text{NH}_3 (\text{aq}) \rightarrow \text{Hg} (\text{s}) + \text{HgNH}_2\text{Cl} (\text{s}) + \text{NH}_4 (\text{aq}) + \text{Cl} (\text{aq})$$
 Would the net be the same as the total. Net Ionic Equation..? | Yahoo Answers I am

having a bit of trouble writing the net ionic equations for these problems. If you can explain how to do them, it would be great! I need to know what the cation and anions are and then also what the product is. I know that some may be soluble and therefore a precipitate is not formed and there is no product, but I am very confused on how to find the equations. net ionic equations? | Yahoo Answers Molecular Equation. $\text{HCl (aq)} + \text{NaHCO}_3 \text{ (aq)} \rightarrow \text{NaCl (aq)} + \text{H}_2\text{O (l)} + \text{CO}_2 \text{ (g)}$ Total Ionic Equation. $\text{H}^+ \text{ (aq)} + \text{Cl}^- \text{ (aq)} + \text{Na}^+ \text{ (aq)} + \text{CO}_3^{2-} \text{ (aq)} \rightarrow \text{Na}^+ \text{ (aq)} + \text{Cl}^- \text{ (aq)} + \text{H}_2\text{O (l)} + \text{CO}_2 \text{ (g)}$... What is net ionic equation for $\text{HCl} + \text{NaHCO}_3$? | Yahoo Answers however, they asked for a net ionic equation. since CO_3^{2-} are on both sides you ignore it in the net ionic equation. 2Ag^+

+ 2 NH₃ + H₂O --> 2NH₄⁺ + 2Ag + OH⁻-Note: ur right Cl⁻ and SO₄²⁻ are the same, since they precipitate. Net Ionic Equations? | Yahoo Answers A solution contains one or more of the following ions: Ag⁺, Ca²⁺, and Co²⁺. Lithium bromide is added to the solution and no precipitate forms. An excess of lithium sulfate is then added to the solution and a precipitate forms. The precipitate is filtered off and lithium phosphate is added to the remaining solution, producing a precipitate. Write a net ionic equation for the formation of the ... net ionic equation help? | Yahoo Answers ionic Sr²⁺ + 2 Cl⁻ + Mg²⁺ + SO₄²⁻ → SrSO₄(s) + Mg²⁺ + 2Cl⁻-magnesium ions and chloride ions are spectators (unchanged) net ionic Sr²⁺ +

$\text{SO}_4^{2-} \rightarrow \text{SrSO}_4(\text{s})$ When you get a good response, please consider giving a best answer. This is the only reward we get. You may have to wait an hour to award BA.

Net ionic equation help!!? | Yahoo Answers

3. net reaction $\text{OH}^- + \text{H}^+ = \text{HOH}$ or H_2O water does not precipitate but does actually form from its ions producing a non ionized product.

4. $\text{Zn}^{2+}(\text{aq}) + \text{S}^{2-}(\text{aq}) = \text{ZnS}(\text{s})$ sodium bromide is soluble so its ions remain separate and are left out of the net ionic equation.

5. Net Ionic Equations.....? | Yahoo Answers

im stuck on some problems :S any help/explanations would be really appreciated!

1. Predict the products of each precipitation reaction balance the complete equation, then write the net ionic

equation a. $\text{Pb}(\text{NO}_3)_2 + \text{KBr} \rightarrow \text{PbBr}_2 + 2\text{KNO}_3$ i got this here.. what should i do from here? *. $\text{Ca}(\text{NO}_3)_2 + \text{KF} \rightarrow$ c. $\text{Ca}(\text{NO}_3)_2 + \text{Na}_2\text{C}_2\text{O}_4$ 2. net ionic equations? | Yahoo Answers Write a balanced total ionic equation and a balanced net ionic equation. From solubility rules, choose three chemicals (NOT: NaBr, NaOH, Na₂S, Na₂SO₄, KIO₃, KClO₃, BaCl₂, NH₄Cl, Pb(NO₃)₂) that could be used for a three bottle experiment. A) Show the balanced molecular equations for your three chemicals. B) Create a grid table showing the interactions and how each chemical could be ... Ionic and Net Ionic Equations? | Yahoo Answers Can someone please help me! I have to predict the product and their states in this reaction and then find the net

ionic equation. Here are the reactants: (the state is in the parenthesis. aq = aqueous. — = rxn. arrow)
 $\text{Na}_2\text{SO}_4(\text{aq}) + \text{ZnCl}_2(\text{aq})$ I used double replacement then to predict the reactions and then balanced the equation: $\text{Na}_2\text{SO}_4(\text{aq}) + \text{ZnCl}_2(\text{aq}) \rightarrow \text{ZnSO}_4(\text{aq}) + 2\text{NaCl}(\text{aq})$... Write the net ionic equation for the reaction? | Yahoo Answers The total and net ionic equations show the real physical state of every component of the reaction. The compound is a strong electrolyte (a soluble ionic compound or a strong acid), then the compound is separated into the appropriate aqueous ions. Net Ionic Equations?!? | Yahoo Answers The net ionic equation for formation of an aqueous solution of NiI_2 accompanied by evolution of

CO₂ gas via mixing solid NiCO₃ and aqueous hydriodic acid is: a. $2\text{NiCO}_3(\text{s}) + \text{HI}(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g}) + 2\text{Ni}^{2+}(\text{aq})$ b. Net Ionic Equations - Chemistry? | Yahoo Answers Write the balanced net ionic equation for each of these precipitation reactions. Enter the cation in the first box and the anion in the second box, with coefficients (if needed, leave 1s off as they are implied). Put charges in parentheses (charge multiple, if any, followed by sign) and phase in parentheses, e.g., $\text{Al}^{3+}(\text{aq})$ should be entered as $\text{Al}(3+)(\text{aq})$. Write the balanced net ionic equation? | Yahoo Answers write a net ionic equation for the following reaction in aqueous solution: sodium sulfate solution is mixed with barium chloride solution. thank you!!!!

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